

# ON THE USE OF SCANNING ELECTROCHEMICAL MICROSCOPY TO INVESTIGATE GALVANIC CORROSION PROCESSES INVOLVING ZINC AND IRON

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## Abstract

The corrosion processes that take place in an iron-zinc galvanic pair exposed to a chloride solution were investigated by scanning electrochemical microscopy (SECM) using a mercury-coated platinum microelectrode as SECM tip. Both the release of zinc ions and the consumption of dissolved oxygen that occur in separate sites at the iron-zinc galvanic pair were monitored. Zinc ions quantification could be performed by stripping the zinc metal collected at the Hg-coated tip using linear sweep voltammetry.

**Keywords:** scanning electrochemical microscopy; galvanic corrosion; iron, zinc, mercury microelectrodes.

## 1. Introduction

*In situ* information on the chemistry and structure at the metal/electrolyte interface is necessary to know the corrosion behaviour of metals and alloys in aqueous environments, and to develop and test efficient protection strategies for longer corrosion resistance of the materials. Nowadays, there is a quite large number of techniques that can be used for corrosion studies, and the choice of the most suitable one to be employed in the investigation depends on various factors, including the surface condition of the materials, dimensions of films and layers on the metal surfaces, and physicochemical characteristics of the aggressive environment. In this way, structural and chemical composition data, as well as surface imaging have been achieved, though, often, measurement procedures involve *ex situ* operations, thus generally quite different from those met by the materials in their environment. Moreover, time-resolved analysis is very difficult at their best, thus making the dynamics of structure modifications produced by the corrosion processes practically impossible to follow. The advent of scanning electrochemical techniques has opened up new opportunities in this field, and among them, scanning electrochemical microscopy (SECM) is conceiving most prospects due to its versatility to examine, with high spatial resolution, the surface topography and the electrochemical reactivity of many metal substrates subjected to corrosion processes [1-4]. Despite the variety of systems characterized by this technique, major limitations have been found for the investigation of a number of technologically interesting metal materials, such as zinc and magnesium. In fact, direct monitoring of the corresponding metal ions could not be investigated because either their corresponding standard redox potentials occur in the range of hydrogen evolution on platinum microelectrodes, which are usually employed as the SECM probe, or they would require the electroreduction of the metal ions to the zero oxidation state, thus effectively modifying the electrochemical response of the SECM tip. On the other hand, those materials are receiving major attention because of lightweighting in areas of automotive and aerospace applications, or as components for active corrosion protection through galvanic coupling and cathodic protection mechanisms. In fact, zinc has been employed in the galvanizing of steels since long ago, but it is also present in modern coating formulations for metal protection designed for offshore conditions, where zinc-rich primers are widely employed. Additionally, zinc-rich primers provide also good performance in the weldability of metallic structures, thus their use is even experiencing a further growth.

Therefore, the development of operation procedures based on scanning electrochemical techniques, opening ways to investigate zinc reactivity *in situ*, is currently a topic of major interest. This is even more the case of scanning electrochemical microscopy,

since direct monitoring of zinc dissolution has not been possible yet with the use of typical SECM tips [5,6]. The introduction of Hg microelectrodes as SECM tips has opened new possibilities in corrosion research [7]. In fact, Hg-probes offer a wider negative potential range than bare platinum tips allowing for the reduction of zinc ions at the tip to be investigated with low interference from oxygen reduction or hydrogen evolution from aqueous solutions [7]. Hg microelectrodes can be fabricated by electrodeposition of liquid mercury onto Pt microdisks [8-11], and adopt a sphere cap geometry, whose size depends on the plating time. Theory to predict diffusion-controlled steady state currents [12-14] and approach curves at sphere-cap tips for both negative and positive feedback conditions [15] is also available.

The objective of this contribution is to assess the ability of SECM coupled to hemispherical SECM tips for monitoring of galvanic coupling reactions involving zinc dissolution. This study has to be considered as a preliminary work for the investigation of electrochemical reactivity and corrosion reactions at metal substrates coated with zinc-rich primers either from cut edges, or from deliberately operated artificial defects through the polymeric layer, which lead to the direct exposure of the underlying metal to the environment. In this work, proof of concept is illustrated for a model zinc-iron galvanic couple immersed in a naturally-aerated aqueous sodium chloride solution.

## **2. Experimental**

The model iron-zinc bimetallic sample consisted of squares with ca. 1x1 mm<sup>2</sup> from each metal to form the testing metal substrates, and they were embedded into an epoxy resin sleeve, such that only the square end surface forms the testing surface. The separation between them was approximately 1.2 mm. Iron and zinc were connected electrically at the back of the epoxy mount to form a galvanic couple. 99.95% pure zinc and 99.5% pure iron, both supplied as sheets of thickness 1 mm by Goodfellow Materials Ltd (Cambridge, UK), were employed. The mounts with the samples were polished with silicon carbide paper down to 1200 grit, washed thoroughly with Millipore deionised water, dried with acetone and finally surrounded laterally by sellotape, thus creating a small container for the test electrolyte solution.

The platinum microdisks, which served as the substrates for mercury deposition, were prepared by sealing wires 25 µm diameter into glass capillaries, which were produced with a conical shape. Prior to mercury deposition, the exposed microdisks were polished with

graded alumina of different sizes on a polishing microcloth. Deposition of mercury onto the platinum microdisks was carried out under potentiostatic control in 5 mM  $\text{Hg}_2(\text{NO}_3)_2$  + 0.1 M  $\text{KNO}_3$  acidified at pH 1 with  $\text{HNO}_3$  [11]. The mercury microelectrode aspect ratio, described by the parameter  $H = h/a$  (i.e. the sphere-cap height ( $h$ ) to the electrode's basal radius ( $a$ ) ratio) was obtained using the method described in ref. [14]. The  $RG$  parameter [17] (that is the overall tip to disk radius ratio) was equal to 5, as evaluated by fitting theoretical and experimental approach curves obtained above an insulating substrate using oxygen as redox mediator [16]. The electrochemical cell was completed using an  $\text{Ag}/\text{AgCl}/\text{KCl}(\text{sat.})$  reference electrode, and a platinum wire as the counter electrode. Tests were carried out in aqueous 0.1 M  $\text{NaCl}$  solution open to air and left at ambient temperature.

A home built SECM apparatus (previously described in ref. [16]) was employed. Linear sweep voltammetry and chronoamperometry were performed using a PAR 283 potentiostat/galvanostat controlled with a personal computer via the M270 PAR software.

### 3. Results and Discussion

In previous work, concerning the corrosion behavior of isolated zinc and iron wires, it was shown that on each metal surface both the anodic and cathodic sites, where the metal dissolution and the oxygen reduction, respectively, occurred, could be visualized [6]. Galvanic coupling of the iron and zinc wires, separated by a region of insulating epoxy resin, leads to the physical separation of the corrosion half cells. Under these conditions, the anodic process takes place on the zinc metal, while the cathodic reaction happens at the iron surface [6]. The distribution of these reactions was detected *in situ* by scanning the Hg-microelectrode polarized at -1.20 V vs.  $\text{Ag}/\text{AgCl}/\text{KCl}(\text{sat.})$ , where the reduction of both  $\text{Zn}(\text{II})$  species and molecular oxygen occurs [7]. SECM line scans were performed across the surface of zinc and iron by positioning the Hg tip at various distances above the substrate investigated. The position of the SECM tip was initially established by using the approach curves performed above the insulating layer of the epoxy resin using oxygen as redox mediator [16]. Figure 1 displays a family of line scans across the two metals obtained accordingly. The generation of soluble  $\text{Zn}(\text{II})$  species due to the corrosion of the zinc is observed as a large negative current when the tip passed over the zinc sample. That is, the cathodic currents measured at the tip grew significantly when the tip was scanned over the zinc metal. On the other hand, the consumption of oxygen in the cathodic reaction that occurs on the iron specimen was detected from the measurement of depleted faradaic currents at the tip. Zinc dissolution in the galvanic pair was found to be spread, though not

uniformly, over the whole area of the metal, conversely to what happened when isolated zinc was exposed to the same aggressive environment (Figure 2). In the latter case, both the anodic and cathodic reactions responsible for the corrosion process were located over the same metal, whereas in the galvanic couple each of the two reactions occurred on a different metal. In Figure 1, it is also worth noting that the cathodic current above the zinc sample is not homogeneously distributed but increases passing from the region closest to iron towards the center of the sample and then decreases drastically towards the other end of the sample. This feature is related to the tendency of zinc to develop pits on its surface during its corrosion in 0.1 M NaCl, and this result agrees well with previous observations derived from the measurement of ionic currents departing from the metal using the scanning vibrating electrode technique (SVET) [6]. Conversely, the oxygen depletion above the iron surface occurred almost uniformly because no preferential sites for this cathodic reaction are established on this metal, again closely matching SVET observations [6].

Scan lines recorded at different tip-substrate distances, also included in Figure 1, show that the ability of the technique to spatially resolve the highly localized anodic process occurring on the zinc metal is greatly reduced as the tip-substrate distance is increased. This feature is a consequence of the diffusion of the ions from the source at the surface into the bulk electrolyte, and therefore the currents become less intense and more distributed laterally as the microelectrode is moved away from the metal.

The effect of the diffusion of the generated Zn(II) ions into the electrolytic solution from the zinc sample could also be investigated from the stripping of the zinc accumulated in the Hg-electrode during the previous scan line. The corresponding stripping voltammetric waves for their oxidation are shown in Figure 3. They were recorded at different electrode-substrate distances immediately after the measurement of the scan lines. It can be observed that the current signals due to metal stripped from the tip greatly depended on the tip-substrate distance. Therefore the latter is a critical parameter that affects the lateral resolution of the concentration profiles of species generated at a corroding substrate. As for the shape of the stripping voltammograms displayed in Figure 3, which present various peaks at different potentials, it can be attributed to a complex electrode process pathway involving also oxygen present in the solution, which is also reduced within the potential region explored. To confirm that the multiple stripping peaks are related to the oxidation of amalgamated zinc, test experiments were performed in  $\text{Zn}(\text{NO}_3)_2$  aerated solutions in a conventional electrochemical cell. Typical stripping voltammograms thus obtained are shown in Figure 4. As is evident the multiple peaks are also observed in the test solutions, thus confirming that the linear sweep stripping peaks recorded in the SECM measurements (Figure 3) belongs to the metallic zinc

accumulated in the mercury layer during the line scan. The concentration profile of zinc ions diffusing from the zinc surface towards the bulk solution was also estimated using a calibration plot based on the charge involved in the stripping voltammograms [18]. In this way it was established that when the tip was positioned at 50  $\mu\text{m}$  above the zinc sample, the concentration of zinc (II) was 1.7 mM; it decreased down to 0.7 mM at 1000  $\mu\text{m}$ .

#### **4. Conclusions**

Mercury microelectrodes can be employed in scanning electrochemical microscopy to detect both the release of zinc(II) ions and the consumption of dissolved oxygen that occurs in the galvanic corrosion of a zinc-iron pair in aqueous solutions without interference from hydrogen evolution.

Dissolution of reduced zinc metal within the mercury layer avoids problems related to chemical and size changes of the tip due to the metal nucleation and growth process, a phenomenon that is typically encountered on solid SECM tips. And sensitive concentration profiles originated by dissolving zinc species from the metal could be determined by stripping voltammetry at the Hg-microelectrode tip.

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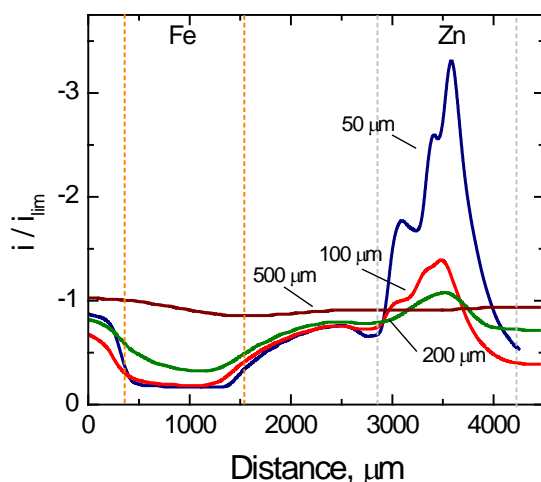
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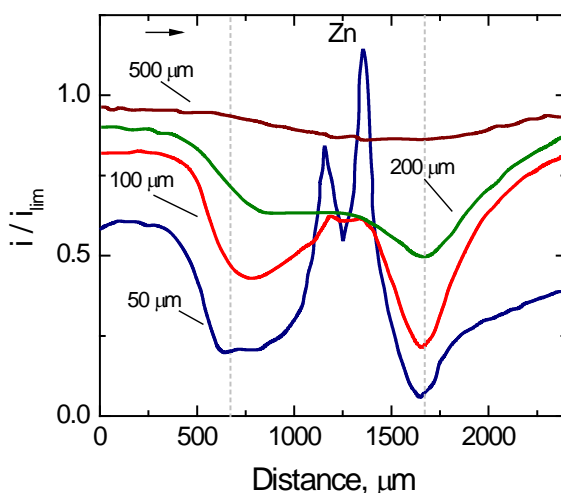
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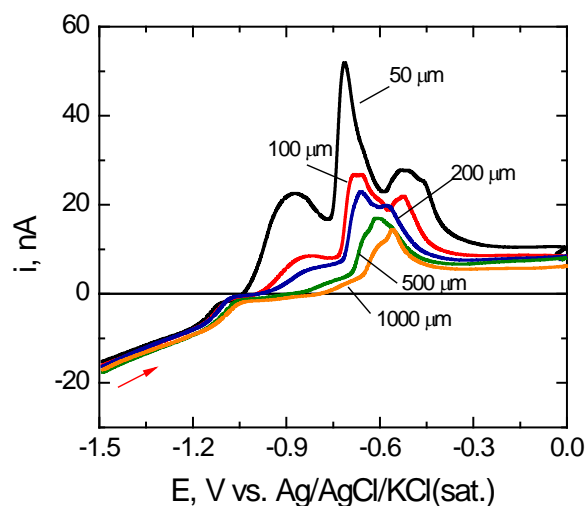


**Figure 1:** Scan lines of a hemispherical mercury tip travelling above an iron-zinc galvanic pair after immersion in 0.1 M NaCl. The plots were measured for different tip-substrate distances as indicated in the graph. Tip potential: -1.20 V vs. Ag/AgCl/KCl(sat.); sphere-cap parameters:  $H = 1$ ,  $RG = 5$ . The arrow indicates the direction of tip movement.

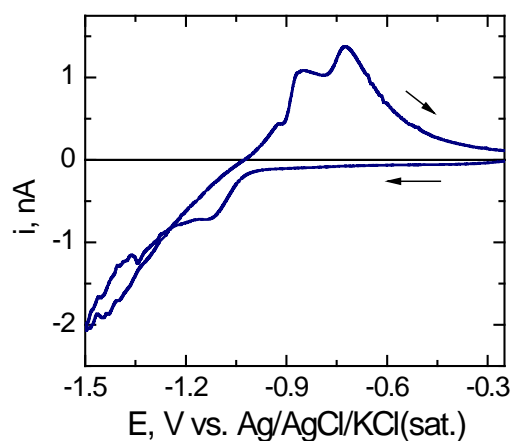


**Figure 2:** Scan lines of a hemispherical mercury tip travelling above a zinc sample after immersion in 0.1 M NaCl. The plots were measured for different tip-substrate distances as indicated in the graph. Tip potential: -1.20 V vs. Ag/AgCl/KCl(sat.); sphere-cap parameters:  $H = 1$ ,  $RG = 5$ . The arrow indicates the direction of tip movement.





**Figure 3:** Anodic stripping voltammograms of the zinc accumulated in the hemispherical mercury tip during the measurement of the scan lines depicted in Figure 1.  $\nu = 0.05 \text{ V s}^{-1}$ . The distances in the plot correspond to the height of the tip above the substrate employed to record the corresponding scan lines over the iron-zinc galvanic pair. Electrolyte: 0.1 M NaCl. Sphere-cap parameters:  $H = 1$ ,  $RG = 5$ .



**Figure 4:** Cyclic voltammogram recorded at a hemispherical mercury tip in the bulk of a 0.1 mM  $\text{Zn}(\text{NO}_3)_2 + 0.1 \text{ M NaCl}$  aerated solution.  $\nu = 0.02 \text{ V s}^{-1}$ .